RAVI MATHS TUITION CENTER, NEAR VILLIVAKKAM RLY STATION, CHENNAI – 82.
WHATSAPP - 8056206308

Alkali and Alkaline Earth Metals 3

11th Standard

Chemistry

Exam Time : 01:15:00 Hrs | Total Marks : 40

20 x 2 = 40

1) Why alkali metals do not occur free in native?
2) Explain the following:
   The second ionisation enthalpy values of alkali metals are high.
3) Mention the uses of lithium.
4) Lithium halides are covalent. Explain.
5) Discuss the biological importance of sodium and potassium.
6) Give methods of preparation of beryllium chloride.
7) How does (i) solubility, (ii) thermal stability, (iii) basic character hydroxides of group 2 elements very down the group.
8) Explain with equations beryllium hydroxide is amphoteric.
9) What happens when carbon dioxide is passed through lime water? Give equation.
10) Explain the term 'desert rose'.
11) What is 'dead burnt plaster'? What is use?
12) Explain the action of hydrogen with alkali metals.
13) Explain the action of sodium with water
14) Atomic radii of alkaline earth metals are smaller than the corresponding members of alkali metals. Why?
15) How would you prepare beryllium hydride from beryllium chloride?
16) Write about the uses of strontium.
17) What is milk of lime? How CO_2 reacts with it?
18) How gypsum occurs in nature?
19) What is meant by setting of cement?
20) Why is Li_2CO_3 decomposed at a lower temperature whereas Na_2CO_3 at higher temperature?

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